

Review Unit 3-Chapter 10,11.1,11.2,12.1	Chem 11- Mrs. Sanford	Name:
Mole-		
Avogadro's number-		
Representative particle means atoms, ions, molecules, formula units, and so on.		
Conversions: Fill in <input type="text"/> \longleftrightarrow mol \longleftrightarrow <input type="text"/>		
Formulas for conversions involving moles \rightarrow $n = \frac{\text{_____}}{M}$ $n = \frac{\text{_____}}{6.02 \times 10^{23}}$		
Note : section 10.2 was done in the gas unit		
Percent composition:		
Formula for % composition \rightarrow		
Empirical Formula \rightarrow give the steps 1) % to mass 2) 3) 4)		
Molecular formula can be drawn with a structural diagram, whereas an empirical formula cannot.		
Molecular Formula-		
In a chemical equation <i>reactants</i> \rightarrow ? Please review symbols at the bottom of p 323		
Catalyst-		
Where is a catalyst written with an equation?		
Coefficient-		
The five types of reactions are (write below) and give an example for each.		
1)		
2)		
3)		
4)		
5)		
Please go over page 338-339 Answer the following:		
For reaction 1, give the another name for the type of reaction example that they give \rightarrow		
For reaction 2, what is the liquid that is formed?		
For reaction 3, what is another way to write water for predicting the products?		
For reaction 4, what is the precipitate?		
For reaction 5, what is burning in the bunsen burner?		
What does mass conservation mean? (see p 357)		

Using Fig 12.3 on p 357 Do #9 on p 358 below:
$\rightarrow 2\text{K(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{KOH(aq)} + \text{H}_2\text{(g)}$
#of particles →
Moles→
Masses(need molar mass first)→
Do p 306 #32, 33, p 307 # 35 p 310 #36,37, p 312 38,39, #45 (also draw a structural diagram of this molecule) Do p315 #52a, 53a,63c, #66(also draw a structural diagram of this molecule using molecular formula)p 316 #73(be careful with this one)p 347 Do 52, p 349 Do #71all